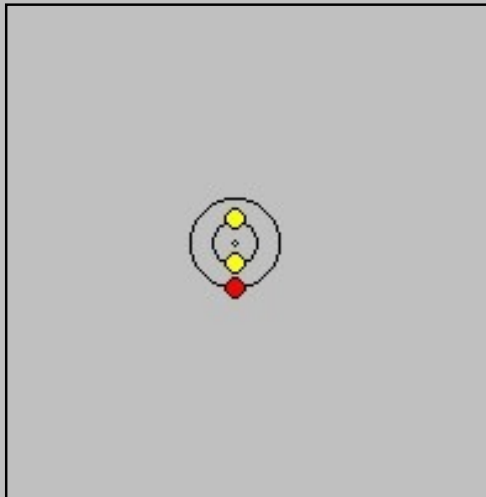




H	Li	Na	K	Rb	Cs	Fr	Uue	Uhe	Bue
1	3	11	19	37	55	87	119	169	219

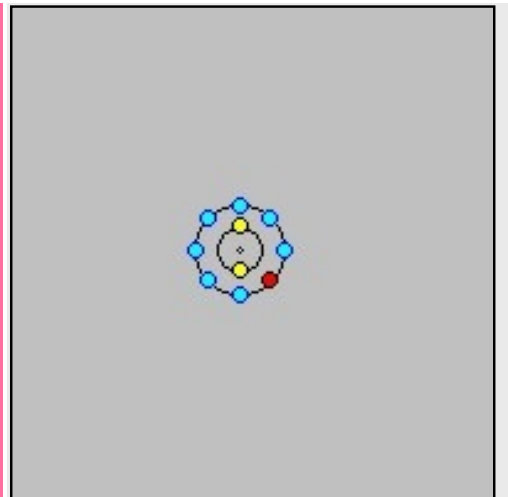
Lithium

Symbol	Li
Atomic Number	3
Relative Atomic Mass $^{12}\text{C} = 12.0000$	6.941
Atomic Radius pm	152
First Ionisation Energy kJ mol^{-1}	513.3
Ionization energy (eV)	5,3917
Electronegativity	0.98
Density kg m^{-3}	534 [293 K] 515 [m.p.]
Molar Volume cm^3	13.00
Thermal Conductivity $\text{W m}^{-1} \text{K}^{-1}$	84.7 [300 K]
Melting Point K	453.69
Boiling Point K	1620
Number of Isotopes	5
Inner + outer Shells	1 + 1 = 2
Inner + outer Orbitals	2 + 1 = 3
Filling Orbital	2s ¹
Ground State Electron Configuration	[He] 2s ¹



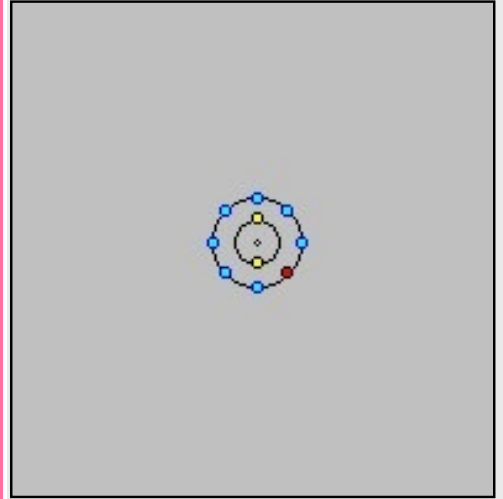
Ground State Electron Configuration with free Orbitals (n= 7)

0, 7



Ground State Electron Configuration with compressed Orbitals (n= 0)

0, 0



Singularity $S(n=2)$

10 = 2 + 1 + 7 + 0

	s	p	d	f	g	h	i	j
1	2							
2	1	1	6					
3								
4								
5								
6								
7								

Term Symbol	$^2S_{1/2}$
Discovery	Discovered by J.A. Arfvedson (Stockholm, Sweden) in 1817 and isolated by W.T. Brande in 1821
Name Derived From	Greek lithos meaning 'stone'