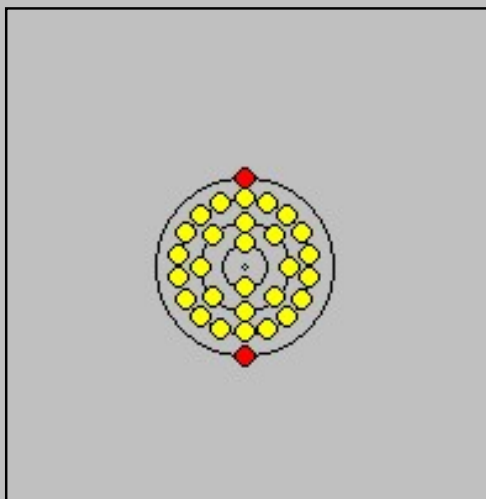




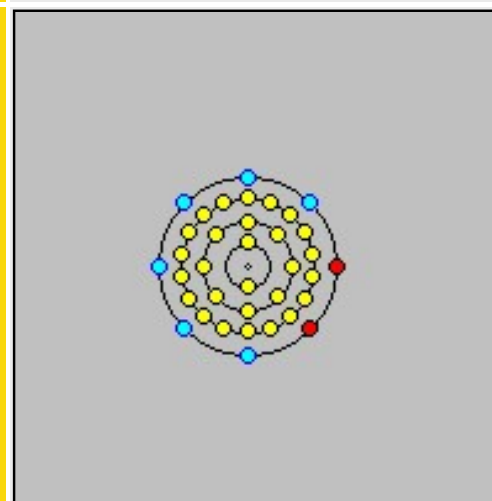
<u>Zn</u>	<u>Cd</u>	<u>Hg</u>	<u>Uub</u>	<u>Uhb</u>	<u>Bub</u>
30	48	80	112	162	212

Zinc

Symbol	Zn
Atomic Number	30
Relative Atomic Mass $^{12}\text{C} = 12.0000$	65.409 (± 4) [Since 2001]
Atomic Radius pm	133
First Ionisation Energy kJ mol^{-1}	906.4
Ionisation Energy (eV)	9.3942
Electronegativity	1.65
Density kg m^{-3}	7133 [293 K] 6577 [l., m.p.]
Molar Volume cm^3	9.17
Thermal Conductivity $\text{W m}^{-1} \text{K}^{-1}$	116 [300 K]
Melting Point K	692.73
Boiling Point K	1180
Number of Isotopes	23
Isotope Atomic mass/u Mole fraction	64Zn 63.929 1461(18) 0.482 68(321) 66Zn 65.926 0364(17) 0.279 75(77) 67Zn 66.927 1305(17) 0.041 02(21) 68Zn 67.924 8473(17) 0.190 24(123) 70Zn 69.925 325(4) 0.006 31(9)
Inner + outer Shells	3 + 1 =4
Inner + outer Orbitals	28 + 2 =30
Filling Orbital	4s ²
Ground State Electron Configuration	[Ar] 3d ¹⁰ 4s ²
Ground State Electron Configuration with free Orbitals (n=6)	
	0, 0, 0, 6

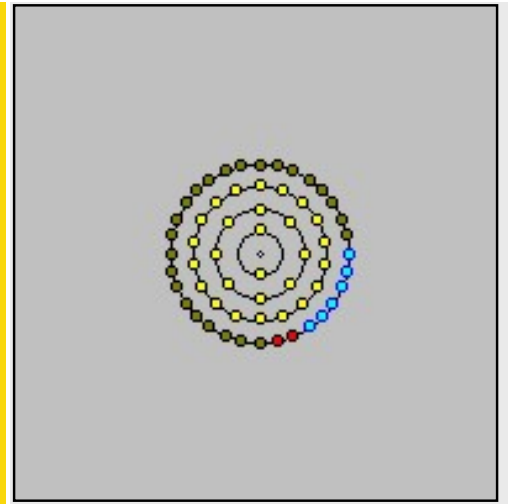


2, 8, 18, 2



Ground State Electron Configuration with compressed Orbitals (n=24)

0, 0, 0, 24



Singularity

60 = 28 + 2 + 6 + 24

	s	p	d	f	g	h	i	j
1	2							
2	2	6						
3	2	6	10					
4	2	6	10	14				
5								
6								
7								

Term Symbol

$1S_0$

Discovery

It was recognized as a metal as early as 1374.

Name Derived From

The name derives from the German zink of unknown origin. It was first used in prehistoric times, where its compounds were used for healing wounds and sore eyes and for making brass.